
Figure 2

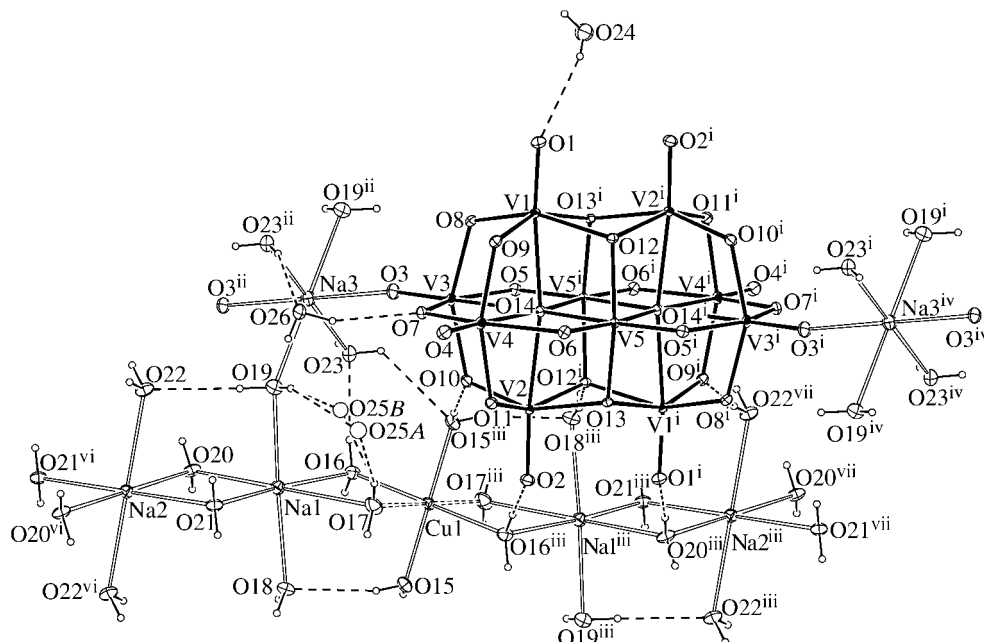
Packing diagram of (I), projected on to the bc plane. Hatched octahedra represent CuO_6 moieties. Open circles denote water O atoms that do not coordinate to Cu^{2+} .

six water molecules, completing discrete octahedra without sharing any ligands with other cations. These aqua ligands accept hydrogen bonds from the aqua ligands of atom Cu1 [O16—H16B \cdots O22 and O15—H15A \cdots O19($-x, 1-y, -z$); Table 2] and donate hydrogen bonds to the surface O atoms of the decavanadate anion.

From solutions containing both Cu^{2+} and Na^+ as cationic species, the decavanadate anion crystallizes with both Cu^{2+} and Na^+ cations to give a double salt, (II). The unit cell of (II) includes a $[\text{V}_{10}\text{O}_{28}]^{6-}$ anion, a Cu^{2+} cation, four Na^+ cations and water molecules. The decavanadate anion, the Cu^{2+} cation and two of the three independent Na^+ cations (Na2 and Na3) are located at the inversion centers. As shown in Fig. 3, the molecular structure of the decavanadate anion in (II) has a metal–oxygen framework identical to that reported by Evans (1966). The Cu^{2+} cation and two of the three independent Na^+ cations (Na1 and Na2) are octahedrally coordinated by six water O atoms. These cations share aqua ligands to form an edge-sharing linear polycationic array of octahedra, *viz.* $\{[\text{Na}1(\text{H}_2\text{O})_2]^+(\mu\text{-H}_2\text{O})_2[\text{Cu}1(\text{H}_2\text{O})_2]^{2+}(\mu\text{-H}_2\text{O})_2[\text{Na}1(\text{H}_2\text{O})_2]^+(\mu\text{-H}_2\text{O})_2[\text{Na}2(\text{H}_2\text{O})_2]^+(\mu\text{-H}_2\text{O})_2\}$

$\text{H}_2\text{O})_2[\text{Na}2(\text{H}_2\text{O})_2]^+(\mu\text{-H}_2\text{O})_2\}_{\infty}$, running along the [011] direction, as illustrated in Fig. 4. The Cu^{2+} cation in (II) also exhibits a Jahn–Teller distortion, with equatorial and axial Cu—O distances of 1.9247 (15)–2.0420 (15) and 2.3896 (18) Å, respectively (Table 3). The water molecules at the axial sites also coordinate to adjacent Na^+ cations, participating in the linkage of the catenary polycation. These longer Cu—O bonds, together with the O15—H15B \cdots O18 and O19—H19B \cdots O22 hydrogen bonds (Table 4), may stabilize the infinite chain. The remaining Na^+ cation, Na3, shares aqua ligands with this array to connect adjacent arrays into a two-dimensional polycationic network, as shown in Fig. 4. Terminal O atoms of the decavanadate anions also coordinate to this Na cation, to form a one-dimensional $\{[\text{Na}(\text{H}_2\text{O})_4]^+\cdots[\text{V}_{10}\text{O}_{28}]^{6-}\cdots\}_{\infty}$ chain structure running along the [111] direction. The decavanadate anions link the polycationic sheet structure into a three-dimensional network.

A similar catenary structure, *viz.* $\{[\text{Na}(\text{H}_2\text{O})_4]^+\cdots[\text{V}_{10}\text{O}_{28}]^{6-}\cdots\}_{\infty}$, has been observed in (HMTA-H \cdots H $_2$ O)-(HMTA-CH $_2$ OH)[H $_3$ V $_{10}$ O $_{28}$ {Na(H $_2$ O) $_4$ }] \cdot 4H $_2$ O (Duraismay *et al.*, 2000; HMTA is hexamethylenetetramine), where hydrogen bonds from the aqua ligands of the Na^+ cations to O atoms of the decavanadate anions support the chain structure. In contrast, the aqua ligands in the $\{[\text{Cu}(\text{H}_2\text{O})_4]^{2+}\cdots[\text{V}_{10}\text{O}_{28}]^{6-}\cdots\}_{\infty}$ chain of (I) and the $\{[\text{Na}(\text{H}_2\text{O})_4]^+\cdots[\text{V}_{10}\text{O}_{28}]^{6-}\cdots\}_{\infty}$ chain of (II) do not form hydrogen bonds to the decavanadate anions. The absence of such hydrogen bonds results in an almost linear Cu/Na \cdots O—V linkage [170.50 (9) $^\circ$ for Cu1 \cdots O3—V3 for (I) and 169.40 (9) $^\circ$ for Na3 \cdots O3—V3 for (II)] as opposed to the bent Na \cdots O—V linkage (141.5 $^\circ$) in (HMTA-H \cdots H $_2$ O)(HMTA-CH $_2$ OH)[H $_3$ V $_{10}$ O $_{28}$ {Na(H $_2$ O) $_4$ }] \cdot 4H $_2$ O. A discrete oligocationic structure, $\{[\text{Na}(\text{H}_2\text{O})_3][\text{Ni}(\text{H}_2\text{O})_6][\text{Na}(\text{H}_2\text{O})_3]\}^{4+}$, has recently been reported in [Ni-


Figure 3

A view of an asymmetric unit and selected neighbors of (II). Displacement ellipsoids are shown at the 50% probability level. The elongated Cu \cdots O contacts are shown as broken lines. [Symmetry codes: (i) $1-x, 1-y, 1-z$; (ii) $-x, -y, -z$; (iii) $-x, 1-y, -z$; (iv) $1+x, 1+y, 1+z$; (vi) $-x, -y, -1-z$; (vii) $x, 1+y, 1+z$.]

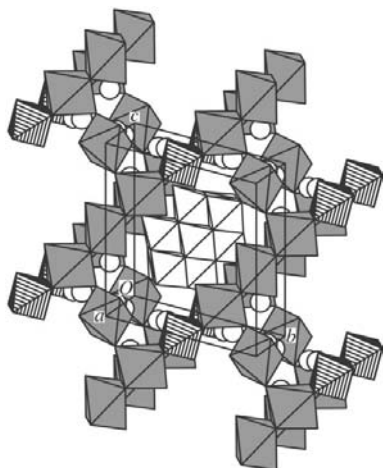


Figure 4

Packing diagram of (II), projected on to the bc plane. Hatched and shaded octahedra represent CuO_6 and NaO_6 moieties, respectively. Open circles represent water O atoms that do not coordinate to any metal cations.

$(\text{H}_2\text{O})_6]_2[\text{Na}(\text{H}_2\text{O})_3]_2[\text{V}_{10}\text{O}_{28}] \cdot 4\text{H}_2\text{O}$ obtained from a similar $\text{Na}^+ - \text{Ni}^{2+} - [\text{V}_{10}\text{O}_{28}]^{6-}$ system (Higami *et al.*, 2002). The difference in the $M^{2+} - \text{O}$ distances {2.053–2.107 Å for Ni^{2+} in $[\text{Ni}(\text{H}_2\text{O})_6]_2[\text{Na}(\text{H}_2\text{O})_3]_2[\text{V}_{10}\text{O}_{28}] \cdot 4\text{H}_2\text{O}$ versus 1.925–2.390 Å for Cu^{2+} in (II)} may lead to different poly/oligocationic structures. Zheng *et al.* (2001) demonstrated that the decavanadate anion works as a template for the self-assembly of Cu^{2+} complexes with organic ligands. Compound (II) exemplifies that the decavanadate anion also acts as a template for the self-assembly of purely inorganic cations.

Experimental

A decavanadate solution was prepared according to the method described by Jahr & Preuss (1965). V_2O_5 (1.48 g) was dissolved in aqueous H_2O_2 [30% aqueous H_2O_2 (20 ml) diluted by water (200 ml)], and the solution was stirred vigorously for 12 h to remove excess H_2O_2 . An aqueous solution (20 ml) containing $\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$ (1.28 g) was added to this solution and the pH was adjusted to 2.98 by the addition of CH_3COOH . The reaction solution was concentrated to a total volume of 175 ml by slow evaporation at room temperature, and the solubility of the target compound was reduced *via* vapor-phase transfer of acetone into this solution. Orange prismatic crystals of (I) were obtained after 2–5 d. To an aqueous solution (50 ml) containing NaVO_3 (0.78 g), $\text{Cu}(\text{CH}_3\text{COO})_2 \cdot \text{H}_2\text{O}$ (0.512 g in 20 ml water) was added, followed by an adjustment of the pH to 3.39 with CH_3COOH . Following vapor-phase diffusion of acetone into this mixture, orange prismatic crystals of (II) precipitated after 1–2 d.

Compound (I)

Crystal data

$\text{Cu}_3\text{V}_{10}\text{O}_{28} \cdot 24\text{H}_2\text{O}$

$M_r = 1580.40$

Triclinic, $P\bar{1}$

$a = 8.8567$ (5) Å

$b = 10.7189$ (6) Å

$c = 11.5492$ (7) Å

$\alpha = 104.662$ (2)°

$\beta = 95.579$ (3)°

$\gamma = 102.064$ (3)°

$V = 1024.14$ (10) Å³

$Z = 1$

$D_x = 2.562$ Mg m⁻³

Mo $K\alpha$ radiation

Cell parameters from 5734 reflections

$\theta = 1.9$ – 27.5 °

$\mu = 3.83$ mm⁻¹

$T = 93$ (2) K

Plate, orange

$0.36 \times 0.23 \times 0.12$ mm

Data collection

Rigaku R-Axis RAPID

diffractometer

Oscillation ω scans

Absorption correction: numerical (Higashi, 1999)

$T_{\min} = 0.359$, $T_{\max} = 0.687$

9090 measured reflections

4688 independent reflections

4486 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.020$

$\theta_{\max} = 27.5$ °

$h = -11 \rightarrow 11$

$k = -13 \rightarrow 13$

$l = -14 \rightarrow 14$

Refinement

Refinement on F^2

$R[F^2 > 2\sigma(F^2)] = 0.022$

$wR(F^2) = 0.077$

$S = 1.34$

4688 reflections

395 parameters

All H-atom parameters refined

$w = 1/[\sigma^2(F_o^2) + (0.0468P)^2 + 0.0054P]$

where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\max} = 0.001$

$\Delta\rho_{\max} = 1.23$ e Å⁻³

$\Delta\rho_{\min} = -0.97$ e Å⁻³

Extinction correction: *SHELXL97*

Extinction coefficient: 0.0433 (12)

Table 1

Selected interatomic distances (Å) for (I).

V1—O1	1.6054 (13)	V4—O9	1.8587 (13)
V1—O8	1.8193 (13)	V4—O11	1.8825 (13)
V1—O9	1.8238 (13)	V4—O6	2.0526 (13)
V1—O12	1.9972 (13)	V4—O14	2.2916 (13)
V1—O13 ⁱ	2.0016 (13)	V5—O6	1.6833 (13)
V1—O14	2.2331 (12)	V5—O5 ⁱ	1.6930 (14)
V2—O2	1.6050 (13)	V5—O12	1.8930 (13)
V2—O11	1.8067 (13)	V5—O13	1.9375 (13)
V2—O10	1.8333 (13)	V5—O14	2.1122 (13)
V2—O13	1.9927 (13)	V5—O14 ⁱ	2.1299 (13)
V2—O12 ⁱ	2.0211 (13)	Cu1—O15	1.9296 (14)
V2—O14	2.2379 (12)	Cu1—O16	1.9690 (14)
V3—O3	1.6041 (14)	Cu1—O3	2.3254 (13)
V3—O7	1.8218 (13)	Cu2—O17	1.9323 (14)
V3—O8	1.8719 (14)	Cu2—O18	2.0166 (14)
V3—O10	1.8903 (13)	Cu2—O19	2.2918 (14)
V3—O5	2.0349 (13)	Cu3—O20	1.9624 (14)
V3—O14	2.3021 (13)	Cu3—O21	2.0018 (15)
V4—O4	1.5992 (13)	Cu3—O22	2.3272 (14)
V4—O7	1.8360 (13)		

Symmetry code: (i) $1 - x, 1 - y, 1 - z$.

Table 2

Hydrogen-bonding geometry (Å, °) for (I).

$D-H \cdots A$	$D-H$	$H \cdots A$	$D \cdots A$	$D-H \cdots A$
O15—H15A \cdots O19 ⁱⁱⁱ	0.83 (3)	1.85 (3)	2.640 (2)	160 (3)
O15—H15B \cdots O24 ^{viii}	0.69 (3)	2.02 (3)	2.694 (2)	165 (3)
O16—H16A \cdots O23 ^x	0.72 (3)	2.00 (3)	2.713 (2)	169 (3)
O16—H16B \cdots O22	0.65 (3)	2.23 (4)	2.881 (2)	173 (4)
O17—H17A \cdots O10	0.72 (3)	1.94 (3)	2.653 (2)	167 (3)
O17—H17B \cdots O4 ^{viii}	0.72 (3)	2.04 (3)	2.726 (2)	161 (3)
O18—H18A \cdots O24 ^x	0.75 (3)	2.08 (3)	2.823 (2)	168 (3)
O18—H18B \cdots O11 ^x	0.72 (3)	2.04 (3)	2.7623 (19)	176 (3)
O19—H19A \cdots O23 ^{viii}	0.73 (3)	2.02 (3)	2.729 (2)	165 (3)
O19—H19B \cdots O26 ⁱ	0.85 (3)	1.84 (3)	2.685 (2)	172 (3)
O20—H20A \cdots O26 ^{xi}	0.74 (4)	2.15 (3)	2.805 (2)	148 (3)
O20—H20A \cdots O1 ^{xi}	0.74 (4)	2.64 (3)	3.095 (2)	122 (3)
O20—H20B \cdots O8	0.73 (3)	2.09 (3)	2.778 (2)	159 (3)
O21—H21A \cdots O13 ⁱ	0.69 (4)	2.01 (4)	2.6868 (19)	170 (4)
O21—H21B \cdots O25	0.81 (4)	2.09 (4)	2.894 (2)	175 (3)
O22—H22A \cdots O9 ^{viii}	0.72 (4)	2.07 (4)	2.786 (2)	172 (3)
O22—H22B \cdots O5	0.73 (3)	2.05 (3)	2.777 (2)	176 (3)
O23—H23A \cdots O10 ^x	0.72 (3)	2.12 (4)	2.820 (2)	166 (3)
O23—H23B \cdots O2	0.72 (3)	2.05 (4)	2.759 (2)	170 (4)
O24—H24A \cdots O25 ^{xii}	0.86 (3)	2.00 (3)	2.838 (2)	167 (3)
O24—H24B \cdots O7	0.78 (4)	2.08 (4)	2.807 (2)	156 (3)
O25—H25A \cdots O1	0.60 (3)	2.17 (4)	2.761 (2)	169 (4)
O25—H25B \cdots O9 ^{xi}	0.79 (4)	2.27 (4)	3.014 (2)	156 (4)
O26—H26A \cdots O12	0.76 (4)	2.04 (4)	2.780 (2)	167 (4)
O26—H26B \cdots O6 ^{xiii}	0.75 (3)	2.22 (3)	2.956 (2)	168 (3)

Symmetry codes: (i) $1 - x, 1 - y, 1 - z$; (iii) $-x, 1 - y, -z$; (viii) $x - 1, y, z$; (ix) $x - 1, y - 1, z$; (x) $1 - x, 1 - y, -z$; (xi) $1 - x, -y, 1 - z$; (xii) $x, y, z - 1$; (xiii) $2 - x, 1 - y, 1 - z$.

Compound (II)

Crystal data

CuNa₄V₁₀O₂₈·23H₂O
M_r = 1527.27
 Triclinic, *P* $\bar{1}$
a = 8.9829 (6) Å
b = 10.7777 (5) Å
c = 11.9761 (9) Å
 α = 104.979 (4)°
 β = 99.028 (2)°
 γ = 99.963 (3)°
V = 1078.12 (12) Å³

Z = 1
D_x = 2.352 Mg m⁻³
 Mo *K*α radiation
 Cell parameters from 4086 reflections
 θ = 1.8–27.4°
 μ = 2.72 mm⁻¹
T = 93 (2) K
 Pillar, orange
 0.20 × 0.14 × 0.07 mm

Table 3

Selected interatomic distances (Å) for (II).

V1—O1	1.6148 (14)	V4—O14	2.2922 (14)
V1—O8	1.8152 (14)	V5—O5 ⁱ	1.6865 (14)
V1—O9	1.8298 (14)	V5—O6	1.6936 (14)
V1—O13 ⁱ	1.9879 (14)	V5—O12	1.9060 (14)
V1—O12	2.0049 (14)	V5—O13	1.9248 (14)
V1—O14	2.2417 (13)	V5—O14 ⁱ	2.1079 (13)
V2—O2	1.6223 (14)	V5—O14	2.1282 (13)
V2—O11	1.7969 (14)	Cu1—O15	1.9247 (15)
V2—O10	1.8239 (14)	Cu1—O16	2.0420 (15)
V2—O13	1.9909 (13)	Cu1—O17	2.3896 (18)
V2—O12 ⁱ	2.0302 (14)	Na1—O20	2.3511 (17)
V2—O14	2.2436 (13)	Na1—O21	2.3565 (17)
V3—O3	1.6014 (14)	Na1—O18	2.3603 (17)
V3—O7	1.8253 (14)	Na1—O19	2.4239 (18)
V3—O8	1.8655 (14)	Na1—O17	2.4553 (18)
V3—O10	1.8755 (14)	Na1—O16	2.4842 (17)
V3—O5	2.0525 (14)	Na2—O21	2.3778 (15)
V3—O14	2.2988 (13)	Na2—O22	2.3888 (15)
V4—O4	1.6029 (14)	Na2—O20	2.3899 (16)
V4—O7	1.8331 (14)	Na3—O3	2.2900 (14)
V4—O9	1.8655 (14)	Na3—O19	2.3765 (17)
V4—O11	1.9024 (14)	Na3—O23	2.4543 (17)
V4—O6	2.0286 (14)		

Symmetry code: (i) 1 - *x*, 1 - *y*, 1 - *z*.

Table 4

Hydrogen-bonding geometry (Å, °) for (II).

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
O15—H15A...O10 ⁱⁱⁱ	0.78 (4)	1.83 (4)	2.600 (2)	170 (3)
O15—H15B...O18	0.71 (3)	2.06 (3)	2.743 (2)	164 (3)
O16—H16A...O23	0.74 (3)	1.94 (3)	2.683 (2)	175 (3)
O16—H16B...O2 ⁱⁱⁱ	0.88 (4)	1.94 (4)	2.820 (2)	174 (3)
O17—H17A...O25A	0.67 (3)	2.01 (3)	2.661 (4)	164 (3)
O17—H17A...O25B	0.67 (3)	2.37 (3)	3.013 (4)	160 (3)
O17—H17B...O11 ^x	0.79 (4)	2.02 (4)	2.811 (2)	174 (3)
O18—H18A...O12 ^{xiv}	0.73 (4)	2.01 (4)	2.729 (2)	170 (3)
O18—H18B...O6 ^x	0.77 (3)	2.06 (3)	2.819 (2)	168 (3)
O19—H19A...O25B	0.77 (3)	1.94 (4)	2.691 (4)	165 (3)
O19—H19A...O25A	0.77 (3)	2.37 (4)	3.076 (4)	154 (3)
O19—H19B...O22	0.80 (4)	2.14 (4)	2.922 (2)	168 (3)
O20—H20A...O1 ^{xiv}	0.68 (3)	2.13 (3)	2.794 (2)	166 (3)
O20—H20B...O8 ⁱⁱ	0.76 (4)	2.04 (4)	2.778 (2)	165 (4)
O21—H21A...O13 ^x	0.76 (3)	1.99 (3)	2.738 (2)	169 (3)
O21—H21B...O24 ^{xii}	0.83 (4)	2.12 (4)	2.941 (2)	171 (3)
O22—H22A...O5 ⁱⁱ	0.79 (3)	2.10 (3)	2.884 (2)	173 (3)
O22—H22B...O9 ^{xv}	0.76 (4)	2.09 (4)	2.844 (2)	173 (3)
O23—H23A...O26 ^{viii}	0.76 (4)	2.13 (4)	2.883 (5)	171 (4)
O23—H23A...O26 ⁱⁱ	0.76 (4)	2.28 (4)	3.015 (5)	162 (4)
O23—H23B...O4 ^{viii}	0.71 (5)	2.09 (5)	2.752 (2)	154 (5)
O23—H23B...O15 ⁱⁱⁱ	0.71 (5)	2.65 (4)	3.137 (2)	127 (4)
O24—H24A...O1	0.72 (5)	2.25 (5)	2.918 (3)	154 (4)
O24—H24B...O9 ^{xi}	0.64 (5)	2.59 (4)	3.118 (2)	143 (5)
O26—H26A...O7	0.80 (7)	2.04 (7)	2.798 (5)	158 (6)
O26—H26B...O24 ^{xii}	0.67 (7)	2.34 (8)	2.980 (4)	161 (8)

Symmetry codes: (ii) -*x*, -*y*, -*z*; (iii) -*x*, 1 - *y*, -*z*; (viii) *x* - 1, *y*, *z*; (x) 1 - *x*, 1 - *y*, -*z*; (xi) 1 - *x*, -*y*, 1 - *z*; (xii) *x*, *y*, *z* - 1; (xiv) *x* - 1, *y*, *z* - 1; (xv) 1 - *x*, -*y*, -*z*.

Data collection

Rigaku R-Axis RAPID diffractometer
 Oscillation ω scans
 Absorption correction: multi-scan (Higashi, 1995)
 T_{\min} = 0.666, T_{\max} = 0.827
 12 066 measured reflections
 4887 independent reflections

4594 reflections with $I > 2\sigma(I)$
 R_{int} = 0.029
 θ_{max} = 27.4°
 h = -11 → 10
 k = -13 → 13
 l = -15 → 15

Refinement

Refinement on F^2
 $R[F^2 > 2\sigma(F^2)]$ = 0.028
 $wR(F^2)$ = 0.068
 S = 1.24
 4887 reflections
 395 parameters
 All H-atom parameters refined
 $w = 1/[\sigma^2(F_o^2) + (0.0307P)^2 + 0.881P]$
 where $P = (F_o^2 + 2F_c^2)/3$

$(\Delta/\sigma)_{\text{max}}$ = 0.001
 $\Delta\rho_{\text{max}}$ = 0.47 e Å⁻³
 $\Delta\rho_{\text{min}}$ = -0.54 e Å⁻³
 Extinction correction: *SHELXL97*
 Extinction coefficient: 0.0126 (5)

In (II), the O25 water molecule is disordered over two sites, *viz.* O25A and O25B, with site occupancies of 0.53 and 0.47, respectively. Because of the correlation between the site-occupancy factors and the displacement parameters, these atoms were refined with isotropic displacement parameters and fixed site occupancies. The H atoms attached to atoms O25A and O25B could not be located. Atom O26 is located close to a center of symmetry, and thus the occupancy of this atom was fixed at 0.5.

For both compounds, data collection: *PROCESS-AUTO* (Rigaku Corporation, 1998); cell refinement: *PROCESS-AUTO*; data reduction: *PROCESS-AUTO*; program(s) used to solve structure: *SHELXS97* (Sheldrick, 1997); program(s) used to refine structure: *SHELXL97* (Sheldrick, 1997); molecular graphics: *ATOMS* (Dowty, 1999) and *ORTEPIII* (Burnett & Johnson, 1996); software used to prepare material for publication: *SHELXL97*.

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Supplementary data for this paper are available from the IUCr electronic archives (Reference: BC1014). Services for accessing these data are described at the back of the journal.

References

Burnett, M. N. & Johnson, C. K. (1996). *ORTEPIII*. Report ORNL-6895. Oak Ridge National Laboratory, Tennessee, USA.
 Dowty, E. (1999). *ATOMS*. Shape Software, Kingsport, Tennessee, USA.
 Duraisamy, T., Ramanan, A. & Vittal, J. J. (2000). *Cryst. Eng.* **3**, 237–250.
 Evans, H. T. Jr (1966). *Inorg. Chem.* **5**, 967–977.
 Hedman, B. (1977). *Acta Cryst.* **B33**, 3083–3090.
 Higami, T., Hashimoto, M. & Okeya, S. (2002). *Acta Cryst.* **C58**, i144–i146.
 Higashi, T. (1995). *ABSCOR*. Rigaku Corporation, Tokyo, Japan.
 Higashi, T. (1999). *NUMABS*. Rigaku Corporation, Tokyo, Japan.
 Jahr, K. F. & Preuss, F. (1965). *Chem. Ber.* **98**, 3297–3302.
 Rigaku Corporation (1998). *PROCESS-AUTO*. Rigaku Corporation, Tokyo, Japan.
 Sheldrick, G. M. (1997). *SHELXS97* and *SHELXL97*. University of Göttingen, Germany.
 Zheng, L.-M., Wang, Y., Wang, X., Korp, J. D. & Jacobson, A. J. (2001). *Inorg. Chem.* **40**, 1380–1385.